

Please find below a list of the main incompatible substances, by way of a non-limiting example.

▶ Acetylene	with copper (piping), halogens, silver, mercury and their compounds
▶ Acetone	with concentrated mixtures of sulphuric and nitric acid
▶ Acetic acid	with chromic acid, nitric acid, hydroxyls, ethylene glycol, perchloric acid, peroxides and permanganates
▶ Chromic acid	with acetic acid, naphthalene, camphor, alcohol, glycerol, turpentine and inflammable fluids
▶ Nitric acid	with acetic, chromic and cyanogenic acid, aniline, carbon, hydrogen sulphide, fluids, gases and substances that are promptly nitrated
▶ Oxalic acid	with silver and mercury
▶ Perchloric acid	with acetic anhydride, bismuth and its alloys, alcohol, paper, wood, fats and other organic substances
▶ Hydrogen sulphide	with nitric acid and oxidants.
▶ Sulphuric acid	with chlorates, perchlorates, permanganates and water.
▶ Alcohols and Polyols	with nitric acid.
▶ Anhydrous ammonia	with mercury, halogens, calcium hypochlorite and hydrogen fluoride
▶ Ammonium nitrate	with acids, metal powders, sulphur, combustible materials
▶ Aniline	with nitric acid and hydrogen peroxide
▶ Silver	with acetylene, oxalic acid, tartaric acid and ammoniac compounds
▶ Arsenic (materials containing it)	with any reducing agent
▶ Azidos	with water.
▶ Chlorine dioxide	with ammonia, methane, phosphine, hydrogen sulphide
▶ Bromine	with ammonia, acetylene, butadiene, butane, hydrogen, sodium carbide, turpentine and finely pulverized metals
▶ Activated carbon	with all oxidizing agents, calcium hypochlorite
▶ Cyanides	with acids and alkali
▶ Chlorates	with ammonia salts, acids, metal powders, sulphur, finely pulverized organic and flammable compounds and carbon
▶ Chlorine	with ammonia, acetylene, butadiene, petrol and other by-products of oil, hydrogen, sodium carbide, turpentine and finely pulverized metals
▶ Chloroform	with sodium and potassium
▶ Chlorides	with sulphuric acid
▶ Dichloromethane	with sodium and potassium
▶ Chlorine dioxide	with ammonia, methane, phosphine, hydrogen sulphide
▶ Fluorine	with all other chemical substances
▶ (White) phosphorus	with air, oxygen, alkali, reducing agents
▶ Hydrocarbons in general	with fluorine, chlorine, formic acid, chromic acid, sodium peroxide
▶ Hydrogen sulphate	with nitric acid vapours and oxidizing gasses
▶ Iodine	with acetylene and ammonia
▶ Hypochlorite	with acids, activated carbon
▶ Flammable fluids	with ammonium nitrate, chromic acid, hydrogen peroxide, nitric acid, sodium peroxide and halogens
▶ Mercury	with acetylene, fulminic acid, hydrogen
▶ Alkaline metals (e.g. calcium potassium, sodium)	with water, carbon dioxide, carbon tetrachloride, and other chlorinated hydrocarbons
▶ Ammonium nitrate	with acids, metal powders, flammable fluids, chlorates, nitrates, sulphur and finely pulverized organic substances or flammable compounds
▶ Nitrites and Nitrates	with acids
▶ Nitroparaffin	with inorganic bases, amines
▶ Calcium oxide	with water
▶ Oxygen	with oils, fats, hydrogen, flammable fluids, solids and gasses
▶ Phosphorus pentoxide	with water
▶ Potassium perchlorate	with sulphuric acid and other acids.
▶ Potassium permanganate	with glycerol, ethylene glycol, benzaldehyde and sulphuric acid
▶ Hydrogen peroxide	with chromium, copper, iron, most other metals and their salts, flammable fluids and other combustible materials, aniline and nitromethane
▶ Sodium peroxide	with any oxidizable substance, such as methanol, glacial acetic acid, acetic anhydride, benzaldehyde, carbon disulphide, glycerol, ethyl acetate and furfural.
▶ Potassium	with carbon tetrachloride, carbon dioxide, water, chloroform, dichloromethane
▶ Copper	with acetylene, azide and hydrogen peroxide
▶ Sodium	with carbon tetrachloride, carbon dioxide, water, chloroform, dichloromethane
▶ Sodium azide	with lead, copper and other metals. This compound is usually employed as a preservative, but it forms unstable and explosive compounds with metals
▶ Selenium	with reducing agents
▶ Sulphides	with strong acids
▶ Carbon tetrachloride	Sodium, potassium